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STANDARD

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JIS K 0102-1 : 2023

**Testing methods for industrial water and
industrial wastewater — Part 1: Test
methods for general physics and
chemistries**
(Amendment 1)

JIS K 0102-1: 2021 was revised under date of 20 July 2023.
This Amendment includes the revised items and is to be used
in conjunction with **JIS K 0102-1: 2021**.

ICS 13.060.25;13.060.30;13.060.50

Reference number : JIS K 0102-1 : 2023 (E)

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Foreword

This Japanese Industrial Standard has been revised by the Minister of Economy, Trade and Industry, through deliberations at the Japanese Industrial Standards Committee in accordance with the Industrial Standardization Act. This Amendment partially replaces the previous edition (**JIS K 0102-1 : 2021**), which has been technically revised.

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Date of Revision: 2023-07-20

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In the event of any doubts arising as to the contents,
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Testing methods for industrial water and industrial wastewater — Part 1: Test methods for general physics and chemistries

(Amendment 1)

JIS K 0102-1: 2021 is revised as follows.

17.2.4 b)

Replace with the following.

b) Carry out one of the following procedures 1) to 3).

1) Add 10 ml of sulfuric acid (1+2) to the sample, and while shaking to mix, add 5 ml of silver nitrate solution (200 g/L). Instead of silver nitrate solution (200 g/L), either of the following may be added:

- 1 g of silver nitrate powder and 5 ml of water;
- 1 g of silver (I) sulfate powder and 5 ml of water.

2) If the sample contains chloride ion, add 10 ml of sulfuric acid (1+2) to the sample, and while shaking to mix, add a volume of silver nitrate solution (200 g/L) that is 5 ml in excess of the weight equivalent to chloride ion. Instead of silver nitrate solution (200 g/L), either of the following may be added:

- an amount of silver nitrate powder that is 1 g in excess of the weight equivalent to chloride ion, followed by 5 ml of water;
- an amount of silver (I) sulfate powder that is 10 % in excess of the weight equivalent to chloride ion, followed by another 1 g of the same powder and 5 ml of water, stirred for 20 min.

NOTE 1 Since silver (I) sulfate is poorly soluble in water, stirring is required to achieve a thorough reaction with chloride ion that reaches the inside of the silver (I) sulfate.

3) If, due to the presence of a large amount of chloride ion in the sample, 10 ml or more of silver nitrate solution (200 g/L) is required, silver nitrate solution (500 g/L) shall be used instead. Add 10 ml of sulfuric acid (1+2) to the sample, and while shaking to mix, add a volume of silver nitrate solution (500 g/L) that is 2 ml in excess of the weight equivalent to chloride ion. Instead of silver nitrate solution (500 g/L), either of the following may be added:

- an amount of silver nitrate powder that is 1 g in excess of the weight equivalent to chloride ion, followed by 5 ml of water;

NOTE 2 The weight of silver nitrate (AgNO_3) equivalent to 1 g of chloride ion is 4.8 g. In the case of sample rich in chloride ion, for example